



## Solar Insolation and Heat Transfer In The Earth

Earth Science/Geology

Mr. Traeger

5. **Geology: Calculate Energy Required to Heat a Substance:** The specific heat capacity of a material is the amount of heat (in Joules) required to raise the temperature of a one-gram mass by 1°C. The formula is:

$$\text{Heat energy (q)} = \text{Mass (m)} \times \text{Specific Heat (c}_s\text{)} \times \text{Change in Temperature ( } \Delta T \text{)}$$

Or

$$q = mc_s \Delta T$$

Calculate the amount of energy (in Joules) that is needed to heat 1 gram of each of the following substances by 10° Celsius:

**Sand (SiO<sub>2</sub>)**

$$c_s = 0.739 \text{ J / } ^\circ\text{C} \times \text{g}$$

**Water (H<sub>2</sub>O)**

$$c_s = 4.18 \text{ J / } ^\circ\text{C} \times \text{g}$$

6. Which substance requires more energy to heat it to the same temperature? Why?
7. From what you have seen in this lab, why do you think that it is generally cooler near the ocean during the day? Why is it generally warmer the farther you go inland **during the day**?
8. Would coastal areas be warmer or colder than deserts **at night**? Why?
9. What would happen to the temperature of the sand if it were black?

### **Part 2 Procedure: Intensity of Insolation**

1. Put new water in your beaker. Make sure that the sand and soil have cooled back down to their original temperatures.
2. Move each beaker **at an angle** under the heat lamp, but *do not* turn on the lamp yet. Try to keep your lamp at the same distance from the beakers as in part 1.
3. Repeat the steps as outlined in Part 1 for recording and graphing data.

### **Part 2 Analysis Questions**

1. Look at your second graph and compare it to your first. How did the temperatures change, if at all?
2. What would the length of the day do for the amount of heating of the earth?
3. What would the time of the day do for the amount of heating of the earth?



