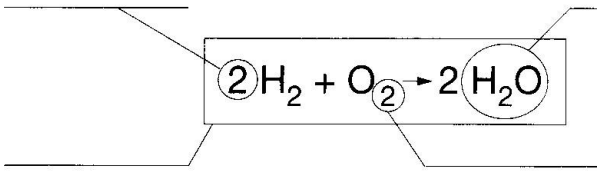
Name: _____ Period: _____

Chemical Formulas

Recall that a compound is two or more elements chemically bonded to one another. Every compound is represented by a **chemical formula** which shows which elements are in the compound. **Subscripts** (small numbers) following each element symbol tell how many atoms of each type are in one molecule of the compound. A **coefficient** in front of the formula tells how many molecules of that compound there are.



Practice: For each of the formulas listed below, tell how many atoms of each element are present in the compound.

A) $NaHCO_3$
Na: _____ H: _____ C: _____ O: _____

B) $C_2H_4O_2$
C: _____ H: _____ O: _____

C) $Mg(OH)_2$
Mg: _____ O: _____ H: _____

D) $3H_3PO_4$
H: _____ P: _____ O: _____

E) $2H_2SO_4$
H: _____ S: _____ O: _____

F) $(NH_4)_3PO_4$
N: _____ H: _____ P: _____ O: _____

G) $C_6H_{12}O_6$
C: _____ H: _____ O: _____

H) $4CaCO_3$
Ca: _____ C: _____ O: _____

Name: _____ Period: _____

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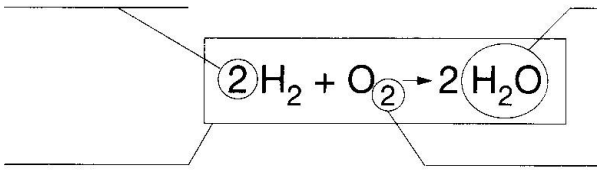

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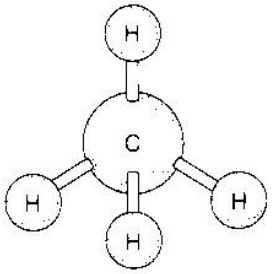
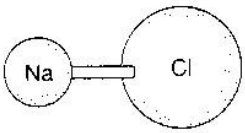
Ca: _____ C: _____ O: _____

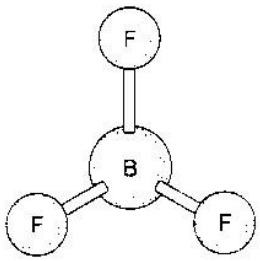
Practice: Under each drawing, write the chemical formula for the molecule shown.



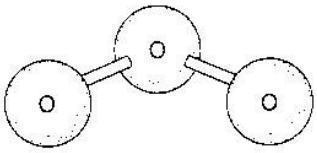
1. _____



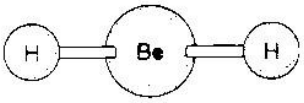
2. _____



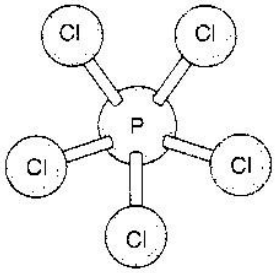
3. _____



4. _____

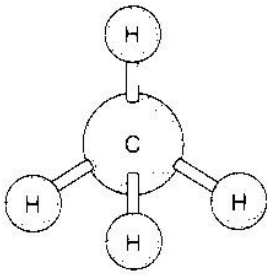


5. _____

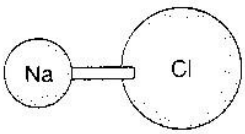


6. _____

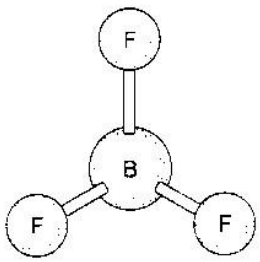
Practice: Under each drawing, write the chemical formula for the molecule shown.



1. _____



2. _____



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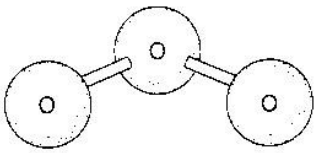
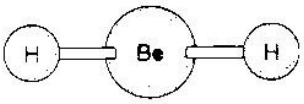
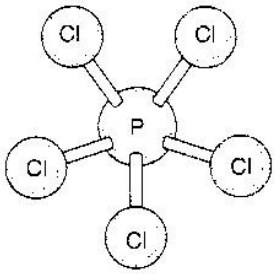


4. _____



5. _____



6. _____