

INSTRUCTIONS FOR TEST CORRECTIONS

Grades for tests may be improved by earning points for missed questions. This is done by demonstrating knowledge of the material that was not mastered at the time of the test. For any question that did not receive full credit, a student may earn a maximum of $\frac{1}{2}$ of the missed points by providing the instructor with an appropriate correction as described below.

- For **each question missed**, submit the following **written** explanation (as 2 separate steps):
 - a. **The number of the question that is wrong and why the answer you gave was incorrect**
 - b. **The correct answer (with complete calculation/explanation included) and why it is preferable to the answer you provided; or if left blank why the answer you show on the corrections is correct.**
- Include all of the corrections for a test on the necessary number of pages of 8.5 inch x 11 inch paper. (typed or NEATLY handwritten).
- Staple the corrections **on top of the original test** WITHOUT changing any of the information on the original and hand it to me in class. If you change your writing on the original, points may not be awarded for correction (because I can't tell where you lost points before!).

I will evaluate your corrections and add any additional points earned to the original test score.

Test corrections are due exactly ONE WEEK after tests are returned. For example, if tests are returned on Monday, the corrections are due the next Monday. Late corrections will NOT be accepted. You must do ALL missed questions to get the maximum possible points back!

Please contact me if you have any questions. Examples have been provided below. If you do not follow the format outlined above you may not be awarded points!

EXAMPLES OF SUITABLE TEST CORRECTIONS

Multiple Choice Question # 3: a. My answer was wrong because I miscalculated the number of neutrons in an atom of zinc-64.
b. The correct answer was b) 34. That answer is correct because the mass number (64) is the sum of protons and neutrons.

Short Answer Question #6: a. My answer was wrong because I forgot to account for the limiting reagent. I said that 10,000 grams of precipitate could be made. I only calculated mass of precipitate formed from sodium chloride, but not from silver nitrate.
b. (Show the correct, completed calculation here!) The correct answer is 0.0537 g. That answer is correct silver nitrate produces the least amount of precipitate (1.391 g from sodium chloride and 0.0537 g from silver nitrate). Silver nitrate is the limiting reagent.

Short Answer Question #8: a. My answer was wrong because I used the incorrect and misspelled name iron chloride.
b. The correct answer is iron (III) perchlorate, because iron is a variable charge metal that requires roman numerals and the anion is actually the polyatomic ion perchlorate.

(Instructions modified from [Chemistry Department at Glendale Community College in Glendale, AZ.](http://web.gccaz.edu/~debis71431/152_Lecture/Test%20Corrections_2.doc))
http://web.gccaz.edu/~debis71431/152_Lecture/Test%20Corrections_2.doc